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Combined Gas Law Worksheet

Combined Gas Law Worksheet #1.

Use the combined gas law to solve the following problems: 1) If I initially have a gas at a pressure of 10.0 atm, a volume of 24.0 liters, and a

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temperature of 200. K, and then I raise the pressure to 14.0 atm and increase the temperature to 300. K, what is the new volume of the gas? 2)

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Solutions. 1) If I initially have 4.0 L of a

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gas at a pressure of 1.1 atm, what will the volume be if I increase the pressure to 3.4 atm? $(1.1 \text{ atm})(4.0 \text{ L}) = (3.4 \text{ atm})(x \text{ L})$ $x = 1.29 \text{ L}$. 2) A toy balloon has an internal pressure of 1.05 atm and a volume of 5.0 L.

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the Combined Gas Law - one of the fundamental Gas Laws. This worksheet contains an explanation of the relationship between the volume of a gas, pressure of a gas and the temperature of the gas. There is an example problem which has been worked out, and 7 problems

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Gas Laws Worksheet #1 - Boyle's, Charles's, Gay-Lussac's, and Combined Gas Law . Solve all problems - you must show your work (including units). The correct answer is

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ideal in parentheses at the end of the problem. Boyle's Law. 1. A gas sample contained in a cylinder equipped with a moveable piston occupied 300.0 mL at a pressure of 2.00 atm.

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Hg = 101.3 kPa = 760 .0 torr Boyle's
Law Problems: 1. If 22.5 L of nitrogen

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at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume? 2. A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L.

Gas Laws Worksheet - New

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Combined Gas Law. The Combined Gas Law combines Charles' Law, Boyle's Law and Gay Lussac's Law. The Combined Gas Law states that a gas' (pressure \times volume)/temperature = constant. Example: A gas at 110kPa at 30.0 °C fills a flexible container with

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Gas Laws (video lessons, examples and solutions)

The combined gas law combines the three gas laws: Boyle's Law, Charles' Law, and Gay-Lussac's Law. It states that the ratio of the product of pressure

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and volume and the absolute temperature of a gas is equal to a constant. When Avogadro's law is added to the combined gas law, the ideal gas law results. Unlike the named gas laws, the combined gas law doesn't have an official discoverer.

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Combined Gas Law Definition and
Examples

Combined Gas Law Problems: $1 \text{ atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa}$ $k = 273 + 0^\circ\text{C}$
A gas balloon has a volume of 106.0 liters when the temperature is 45.0°C and the pressure is 740.0 mm of mercury. What will its volume be at

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20.0 °C and 780.0 mm of mercury
pressure?

Gas Laws Worksheet #2: Boyle,
Charles, and Combined Gas Laws
This quiz/worksheet combo will test
your knowledge of the combined gas
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