

Section 111 Describing Chemical Reactions Answer Key

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SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329)

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11.1 Describing Chemical Reactions 11 - Evaluation 2016

Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$...

Chapter 11: Chemical Reactions

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11.1 Describing Chemical Reactions 11

Section 11.111.1. Objectives. Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabularyreactant: the starting substance in a chemical reaction. New Vocabularystoichiometrymole ratio

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The term oxidation was originally used to describe chemical reactions involving O₂, but its meaning has evolved to refer to a broad and important reaction class known as oxidation-reduction (redox) reactions. A few examples of such reactions will be used to develop a clear picture of this classification.

4.2 Classifying Chemical Reactions - Chemistry 2e | OpenStax

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