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$S(l) + O_2(g) \rightarrow SO_2(g)$. Yield and rate are increased in this step by providing plenty of oxygen. The second step is heating sulfur dioxide in oxygen to convert it to sulfur trioxide (using a catalyst, vanadium): $2SO_2(g) + O_2(g) \rightarrow 2SO_3(g)$. Yield and rate are increased here by passing over several catalyst beds.

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