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Graham's Law. • speed of diffusion of 2 different gases is inversely proportionality to the square root of their densities. Raoult's Law. • vapor pressure lowering due to effects of electrolytes or solutes in solution. • the partial vapor pressure of the solvent of an ideal solution is equal to the product of the mole fraction of the solvent and its vapor pressure in pure state.

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Dr Valizadeh (Teaching Materials) - Pharmacy faculty

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Objectives of the lecture After completion of this lecture, the student should be able to: 1. Understand the various types of pharmaceutical solutions. 2. Define solubility, saturated & unsaturated solutions and polar & non polar solvents. 3. Describe complete & partial miscibility. 4.

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- Physical properties - a characteristic that can be observed or measured without changing the identity or composition of the substance
- Physical properties used to describe matter can be classified as: 1) Extensive – depends on the . amount. of matter in the sample - e.g. Mass, volume, length . 2) Intensive – depends on the . type. of

THE PROPERTIES AND STRUCTURE OF MATTER

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