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Smith, Jr., Ph.D in Chemical Engineering,
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50 - 200. All substances are in a supercritical state above the critical point, but those with high critical pressure and critical temperature are not practical.

Carbon dioxide is a commonly used supercritical fluid since its critical temperature is just 31.1 degrees C and its critical pressure is only 7.38 MPa.

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A supercritical fluid is the phase of a material somewhere between a gas and a liquid, a dense gas. It happens in the area on a phase diagram where the boundaries between liquid and gas become blurred. When a material is described as supercritical, it stops behaving as a normal liquid or gas. Increasing the temperature of

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a supercritical fluid cannot change it into a gas, whilst increasing the pressure doesn't result in a liquid forming. As you change you would expect with normal liquids and ...

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properties in the pseudo-critical region of supercritical fluids bring very big challenges to the traditional Dittus-Boelter-type heat transfer correlations. In this chapter, we will talk about the principles and applications of two kinds of heat transfer correlations of supercritical fluids: the empirical type and the semiempirical

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