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During the first year of high school chemistry or the first semester of college chemistry, all the terms, units of measurement, and atoms, molecules, elements, and compounds may seem a bit overwhelming. The good news: Your Chem I class doesn't have to be torture. This handy Cheat Sheet provides some basic information you can refer [...]

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The AP Chemistry course is taught as a continuation of the first year course. As such, units are numbered accordingly. Units 1-8 are covered during the first year course and reviewed over a short time span during the AP course. (Links for these units can be found under the "CHEMISTRY" tab above and on the links immediately below this overview.)

[Units 1-5 Pre-AP Chemistry Review - Mrs. Forest's ...](#)

Chemistry Unit Review Sheet- Part 2. Read the descriptions below and determine whether each item describes an endothermic or an exothermic reaction. _____ 1. A reaction that

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produces heat energy _____ 2. A reaction that absorbs energy _____ 3. Paper burns _____ 4

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Unit 5 Review. Show all questions <= ... 2.5 moles per second; Given the following reaction: $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2 + 4H_2O(g)$ If the initial rate of the reaction produces 2 moles of water vapor, H_2O , per second, what is the rate at which propane, C_3H_8 , is consumed initially? ?

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$36.5 \text{ g HCl} = \text{moles of solute} \times \text{liters of solution}$
 $0.200 \text{ M} = \frac{\text{mol solute}}{0.250 \text{ L}}$
 $\text{mol solute} = 0.0500 \text{ mol HCl}$
 $\text{g HCl} = 0.0500 \text{ mol HCl} \times 36.5 \text{ g HCl/mol HCl} = 1.83 \text{ g HCl}$
A solution of Na_2CO_3 contains 65.0 g of solute dissolved in water to make a 3.00 M solution. What is the volume of the solution, in liters? ?
 $\text{g HCl} = 0.0500 \text{ mol HCl} \times 36.5 \text{ g HCl/mol HCl} = 1.83 \dots$

[10-24a,b,c,d Unit 10 Review Sheet-Key](#)

Chemistry 12 Unit 1 - Reaction Kinetics Unit 1 - Review Sheet Page 2 2. For each of the

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following reactions find a quantity or property which could be monitored in order to measure the rate of reaction See p. 2-5 in SW. ("a" is done as an example.) a) $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ pressure will decrease as reaction proceeds because

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