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Table Explained: Introduction

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01 - What Is Oxidation? Learn the
Definition of Oxidation, Oxidation
Numbers & Oxidizing Agents
Balancing Chemical Equations

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1 0.0006 mm = m Answer 2 0.054 mL =
nL

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Chemistry 11 Course Review Chemistry
11 Course Review Page 8 2. The density
of liquid ethanol (C_2H_5OH) is 0.790
 g/mL . Calculate the number of molecules
in a 35.0 mL sample of liquid ethanol.

(NOTE: You CAN'T use 22.4 L/mol since
this is NOT a gas at STP!) Answer _____

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3. Calculate the density of $\text{PCl}_3(\text{g})$ at STP.

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Review. Density 1. A 0.0460 L piece of
copper has a mass of 410.32 g. Calculate

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the density of copper in g/mL. Answer _____

2. Give the number of significant digits in each of the following. Assume they are all measurements.

a) 0.0023 _____

d) 3.2×10^{-4} _____

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Basics 1. a) Goggles must be worn any
time there is an open flame or hazardous

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chemicals in the lab. b) Chemicals should be handled with care; lids should be returned when a bottle is finished; chemicals should never be eaten or tasted unless directed by the teacher.

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Chemistry 11 Course Review Chemistry
11 Course Review Page 5 3. Review
p.53-58. Answer the following: a) Explain
how distillation can be used to separate the
substances in a solution. b) What types of
mixtures does paper chromatography work
best for? c) What is the simplest, most
economical method of separating

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suspensions?

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11. meteorites. Chapter Review. Part A.
Vocabulary Review (page 135) 1. eclipse
(3/1) 2. solar system (4/2) 3. rotation (1/1)
4. revolution (1/1) 5. constellation (7/3) 6.
galaxy (8/3) Part B. Concept Review (page
135) 1. Answers will vary, but should
mention Earth's rotation and that the

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Sun's light falls on only half of Earth at one ...

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include: A course framework: the framework defines what students should know and be able to do by the end of the course. It serves as an anchor for model lessons and assessments, and it is the primary document teachers can use to align instruction to course content.

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Answer to PreAP Chemistry Unit 11
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Review of the Grade 11 course, following
this there will be an evaluation on the
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Use the periodic table to find the following information concerning the element with 15 protons: a. Get the atomic number 15 b.

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Identify the element Phosphorus c. Find the mass number of the most common isotope of an element 31 d. Get how many neutrons the most common isotope has 16 e. How many valence electrons an element has 5 f. What is its charge/oxidation number for the atom-3 g.

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