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Section 9.2 Classifying
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your textbook, read
about synthesis,

combustion,

decomposition, and

replacement reactions.

Assume that Q, T, X,

and Z are symbols for

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elements. Match each equation in Column A with the reaction type it represents in Column B.

Column A 2. $Q + Z \rightarrow QZ$

Answer the following questions. a. b. c. d.

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In a chemical reaction,
chemical (blank) are
formed or broken
exothermic; product

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During an (blank) reaction, more energy is required to form the bonds in the (blank) than needed to break the bonds of the product are formed.

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Concentration increase,
less surface area, higher
temperature, and a

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catalyst. Four ways a chemical reaction has taken place. Color change, the formation of a precipitate, gas formed, temperature change. 3 examples of a physical change.

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single replacement
reaction. answer. A
chemical change in
which one element
replaces a second
element in a compound.
question. double
replacement reaction.
answer. A chemical
change involving an
exchange of positive

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ions between two
compounds. question.
combustion reaction.

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0 o u u o o 0 o o 0 o o o

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Textbook. Used to show
what a chemical
reaction yields (same as
=); used to separate the
reactants from the

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Reactions. Left side items
are reactants and right
side items are the
product. (g) = gas; (aq)
= water; (l) = liquid.
these symbols used to
denote the state of the
reactants and products
in an accurate balanced
chemical equation.

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1. State the term for the amount of energy that is needed for a chemical reaction to occur. 2.

Summarize the relationship between an enzyme and a substrate. 3. Classify which of the compounds in the reaction below are reactants and which are products. $\text{PbO}_2 + 4\text{HCl} \rightarrow \text{PbCl}_2 + 2\text{H}_2\text{O}$ 4.

Compare endothermic

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reactions and
exothermic reactions. f
5.

Answers

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Please answer the
following questions: 1.
What is a chemical

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: Answer the questions and complete statements from viewing the video. Everything is made of chemicals. When do chemical reactions happen? When the electrons hook together. Water is . two. parts “H” and . one. part “O”.

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What makes iron rust
and our blood red?
Oxygen . Car and
motorcycle engines are
controlled by what
chemical reaction?

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Eastern Shore's ...

Check your
understanding of
chemical reactions and
ways to balance them in

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this quiz and printable worksheet. Topics you will need to know in order to pass the quiz include reactants and chemical ...

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Section 7.1 Energy

Goals To introduce the terms energy, kinetic energy, and potential energy. To introduce the Law of Conservation of Energy. To describe the relationships between stability, capacity to do work, and potential energy. To explain why breaking chemical bonds requires energy and why the formation

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Answers ...

For most chemical
reactions in this book,
solids are labeled (s),

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liquids (l), and gases (g). The numerical coefficients in front of the chemical formulas express the moles of each compound or element. The preceding reaction can be interpreted in terms of moles or masses. (See Table 1.)

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When a chemical reaction occurs, the original substances put together, called reactants, lose their chemical properties and become different substances called products with a different set of chemical properties. Reactions where energy is released are called exothermic

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reactions. When energy
is absorbed, it is called
an endothermic reaction.

Chemical reactions. 8th

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10. Ritter, Plumb, et al;

Nelson 2001. Chemistry

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