

Chapter 9 Section 3 Stoichiometry Answers

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~~Example 2: Solution Stoichiometry—Volume to Volume Conversion. A student takes a precisely measured sample, called an aliquot, of 10.00 mL of a solution of FeCl₃. The student carefully adds 0.1074 M Na₂C₂O₄ until all the Fe³⁺ (aq) has precipitated as Fe₂(C₂O₄)₃ (s). Using a precisely measured tube called a burette, the student finds that 9.04 mL of the Na₂C₂O₄ solution was ...~~

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~~mc06se-cfMsr-i-vi~~

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carefully adds 0.1074 M Na₂C₂O₄ until all the Fe³⁺

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Steps for Stoichiometry: 1- Identify the given and target compound 2-Balance the equation for the reaction 3- Set up the problem (convert to moles if necessary) 4-Use the mole ratio(s) to calculate the number of moles of the desired compound 5- convert to grams of the desired compound if necessary

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