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provided. 1. Given the following

equation: $C_3H_4(g) + xO_2(g) \rightarrow$

$3CO_2(g) + 2H_2O(g)$ _____ a.

What is the value of the

coefficient x in this equation?

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$3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient . x. in this

equation? b. What is the molar

mass of C_3H_4 ? c. How many

moles are in an 8.0 g sample of

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provided. 1. 88% The actual yield
of a reaction is 22 g and the
theoretical yield is 25 g. Calculate
the percentage yield. 2. 6.0 mol
of N_2 are mixed with 12.0 mol of
 H_2 according to the following
equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...

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following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ _____

- a. What is the value of the coefficient x in this equation?
40.07 g/mol h. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O c. What is the mole ratio of O_2 to H_2O in the above equation?
0.20 mol d. How many moles ...

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2 60 mol of N_2 are mixed with 120 mol of H
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