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*Compounds Grade 9 Chemistry,
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Patterns in the Table*

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Compounds P-4 | Valence Bond*

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Chapter 9: Chemical Names and Formulas. STUDY. PLAY. Monatomic Ion. Consists of a single atom with a

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Positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

~~Chapter 9: Chemical Names and~~

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29 terms. hales913. Chapter 9:

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Steele. STUDY. PLAY. Monatomic ion.

- single atom with a positive or negative charge resulting from the loss or gain of electrons, respectively.

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Section 9.4 – Naming and Writing
Formulas for Acids and Bases. An acid
is a compound that produces H^+ ions
when it dissolves in water. The formula
for an acid normally starts with an H.

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When naming acids, you should first determine whether or not the acid contains oxygen.

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Formulas. Acids. base. Cation.

Ionic compounds. Compounds that contain one or more hydrogen atoms and produce.... An ionic compound that produces hydroxide ions when dissolved.... Any atom or group of atoms that has a positive charge.

Compounds composed of cations and

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Chapter 9 Chemical Names and
Formulas83 SECTION 9.3 NAMING
AND WRITING FORMULAS FOR
MOLECULAR COMPOUNDS (pages

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268–270) This section explains the rules for naming and writing formulas for binary molecular compounds.

Naming Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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Name _____ Date _____ Class _____
~~CHEMICAL NAMES
AND FORMULAS 9~~

~~ANSWER KEY~~
SECTION 9.2 NAMING AND

WRITING FORMULAS FOR IONIC

COMPOUNDS I. Write the formulas
for these binary ionic compounds. c.

potassium iodide 2. write the formulas

for the compounds formed c. b. f. 3.

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Name the following binary ionic compounds. sodium sulfide g. CuCl : h. SnCl_2 i. a. MnO_2 c. d. SrBr_2 221 Chapter 9 Chemical Names and Formulas

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Formulas 9.1 Naming Ions 9.2 Naming
and Writing Formulas for Ionic
Compounds 9.3 Naming and Writing
Formulas for Molecular Compounds
9.4 Naming and Writing Formulas for
Acids and Bases 9.5 The Laws
Governing How Compounds Form

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Pittsfield

Chapter 9 “Chemical Names and Formulas” ... in parenthesis after the name of the metal (Table 9.2, p.255)
Predicting Ionic Charges Some of the post-transition elements also have more than one possible oxidation

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state. Tin (II) = Sn^{2+} Lead (II) = Pb^{2+}
Tin (IV) = Sn^{4+} Lead (IV) = Pb^{4+}

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Chemical Names and Formulas 281

CHAPTER 9 Assessment 42. a. 2- b.

1+ c. 1- d. 3+ 43. a. 2+ b. 2+ c. 3+

d. 1+ 44. a. barium ion b. iodide ion c.

silver ion d. mercury(II) ion 45.

cyanide, CN^- and hydroxide, OH^- 46.

a. hydroxide ion b. lead(IV) ion c.

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sulfate ion d. oxide ion 47. zero; A
compound is electrically neutral. 48.
The symbols for the cation and

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Chapter 9 Practice Problems:

Chemical Names and Formulas.

Section 9.1: Naming Ions. 1. What is

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the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr^{3+} c. Li^+ e. Cl^- b. P^{3-} ...

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chemical compound, the masses of
the elements are always in the same
proportions any atom or group of
atoms that has a negative charge ...
place cation name first followed by the
anion name: naming polyatomic ...

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