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Equilibrium chapter 18 reaction rates and  
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of Reaction

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~~Chapter 18 Section 2: Progress of  
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## Chapter 18 - Reaction Rates and Equilibrium

a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products occurs in the chemical

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The rate of a chemical reaction is expressed as the amount of reactant changing per unit time. Work Step by Step

The rate of a chemical reaction is

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## Equilibrium. 18.1 Rates of Reaction.

- The time needed to complete a chemical reaction can vary tremendously.
- Rates measure changes that occur within intervals of time.
- In chemistry, rates or speeds are expressed in chemical terms. Atoms, ions, and molecules must collide in order to react

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molecules can form a chemical bond when they collide, as long as the particles have enough kinetic energy and have the proper orientation.

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describes the change over an interval of  
time. collision theory. atoms, ions, and

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molecules can react to form products when they collide, provided that the particles have enough kinetic energy. activation energy.

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Fraser. How is the rate of a chemical change expressed? in chemistry, the rate of chemical change or the reaction rate is usually expressed as the amount of reactant changing per unit time.

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