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Chapter 16: Solids, Liquids, and Gases
628 CHAPTER 16 LIQUIDS AND SOLIDS 22. NaCl, MgCl₂, NaF, MgF₂, and AlF₃ all have very high melting points indicative of strong intermolecular forces. They are all ionic solids. SiCl₄, SiF₄, F₂, Cl₂, PF₅, and SF₆ are nonpolar covalent molecules. Only LD forces are present.

CHAPTER 16 LIQUIDS AND SOLIDS - UCSB
Solids are very difficult to compress and gases are very easy. Liquids are in the middle, but are difficult. When you compress something, you force the atoms closer together. When the pressure goes up, substances are compressed. Liquids already have their atoms close together, so they are hard to compress.

Liquids - Chapter 16- Solids , Liquids and Gases
478 CHAPTER 16 Solids, Liquids, and Gases Liquid State What happens to a solid when thermal energy or heat is added to it? Think about the ice cube in the hot soup.The particles in the hot soup are moving fast and colliding with the vibrating particles in the ice cube. The collisions of the particles transfer energy from the soup to the ice cube.

Chapter 16: Solids, Liquids, and Gases
Chapter 16- Solids , Liquids and Gases: Created by Jessica Armstrong: Home; Solids; Liquids; Gases; Plasma; Kinetic Theory; Pascal & Bernoulli's Principle; Vocab; Sources; This Website explains what the States of matter , Solid, Liquids and Gases are made of and how they work. It also talks about the Kinetic Theory and the energy's that work ...

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Chapter 16 --Liquids and Solids Liquid Interactions Viscosity: The resistance of a fluid (a gas or a liquid) to flow. The higher the intermolecular forces the the viscosity. Viscosity usually decreases as the temperature rises. Molecules have more energy at higher temperatures and can overcome the

Chapter 16: Phenomena
Physical Science Chapter 16 Solids, Liquids, and Gases. At the _____ the pressure of a liquid's vapor is equal to the pressure of the atmosphere, and liquid becomes a gas. Heating curve of a liquid- as a solid melts and a liquid vaporizes, the temperature remains _____; the temperature will increase after the attractive forces of the earlier state have been overcome.

Physical Science Chapter 16 Solids, Liquids, and Gases ...
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Chapter 16 solids, liquids, and gases. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Moss5. Terms in this set (17) States of matter (phases of matter) A condition in which matter can exist. Examples: Solids, Liquids, gases, and plasmas. solid. A state of matter in which a substance hasa definite shape and volume.

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A solid has a definite size and shape. Atoms in a solid vibrate or jiggle but do not move from place to place. Most solid materials have a specific type of geometric arrangement formed by a solid. Chemical and physical properties of solids often can be attributed to the type of geometric arrangement that the solid forms.

Solids - Chapter 16- Solids , Liquids and Gases
16. Table of Contents. 16. Unit 4: The Nature of Matter. Chapter 16: Solids, Liquids, and Gases. 16.1: Kinetic Theory. 16.2: Properties and Fluids.

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16. Section 1 cont. -u>--THERMAL EXPANSION- increase in the Size of a substance when the temperature increases ---The size of the substance will then decrease when the temperature decreases ---Expansion and contraction occur in most solids liquids and gases -. 17.

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chapter chapter OrganizerOrganizer - sd273.com. chapter chapter OrganizerOrganizer 474A CHAPTER 16 Solids, Liquids, and Gases Chapter Opener Section 1 Kinetic Theory 2 sessions 1 block 1. Explain the kinetic theory of matter. 2. Describe the particle movement in the four states of matter. 3. Explain particle behavior at the melting and boiling points.