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Electron Configurations Part 1-
Electrons and Sublevels Electron
Configuration - Basic introduction
How to Write the Electron
Configuration for an Element in
Each Block Electron Configuration

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C 1s 2s 2p 3s

As an example, the ground state configuration of the sodium atom is $1s^2 2s^2 2p^6 3s^1$, as deduced from the Aufbau principle (see below). The first excited state is obtained by promoting a 3s

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electron to the 3p orbital, to obtain the 1s 2 2s 2 2p 6 3p 1 configuration, abbreviated as the 3p level. Atoms can move from one configuration to ...

Electron configuration - Wikipedia

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This decides the electron capacity of the shells. The K shell contains a 1s subshell hence it can carry 2 electrons, the L shell has 2s and 2p, and can carry 8 electrons. The M shell contains 3s, 3p, and 3d, and can carry 18 electrons. The N shell containing 4s, 4d, 4p and 4f,

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can carry 32 electrons.

Electron Configuration Chart for
All Elements in the ...

However there are numerous
exceptions; for example the
lightest exception is chromium,

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which would be predicted to have the configuration $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4 4s^2$, written as $[\text{Ar}] 3d^4 4s^2$, but whose actual configuration given in the table below is $[\text{Ar}] 3d^5 4s^1$.

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Electron configurations of the elements (data page ...

An atom has the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^5$. The electron dot symbol for this element is? a. X surrounded by seven dots. b. X surrounded by five dots. c. X

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surrounded by two dots. d. X

surrounded by three dots.

Explanation please?? :) Answer

Save. 1 Answer. Relevance.

Genuine. Lv 5. 8 years ago.

An atom has the electron

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configuration $1s^2 2s^2 2p^6$
 $3s^2 \dots$

A. $1s^2 2s^1$. B. $1s^2 2s^2 2p^5$.

C. $1s^2 2s^2 2p^6 3s^2$. D.

$1s^2 2s^2 2p^6 3s^2 3p^1$.

Answer Save. 1 Answer.

Relevance. William. 4 years ago.

Favorite Answer. B (fluorine) B

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needs to gain one more electron to complete its p orbitals and thus its second shell. It can do this by taking an electron from sodium (forming an ionic bond). Sodium will ...

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If sodium (Na) has an electron configuration of $1s^2 2s^2 \dots$

A. $1s^2 2s^2 2p^6 3s^2$ B. $1s^2 2s^2 2p^6 3s^2 3d^4$ C. $1s^2$

$2s^2 \dots$ " in Chemistry if there is no answer or all answers are wrong, use a search bar and try to find the answer among similar

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questions.

Which of the following is a reasonable ground-state ...

The electron configuration of an atom is $1s^2 2s^2 2p^6 3s^2 3p^4$. The number of unpaired electrons in

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this atom are b. 3 c. 5 d. no
correct answer given 22. The
correct electron sublevel
structure for 25Mn is a. 1s 2s 2p
3s 3d b. 1s 2s²2p⁶2d⁶3s 3p c.
1s²2s²2p 3s 3p 4s²4p d. 1s 2s 2p
3s 3p 4s 3d 23.

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Solved: A. 2 21. The Electron Configuration Of An Atom Is ...
As orbitals correspond to number of the subshell. 1 corresponds to s orbital. 2 corresponds to 2s 2p. 3 corresponds to 3s 3p 3d. 4 corresponds to 4s 4p 4d 4f. Thus

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1s 2s 2p 3s 3p 3d the next will be
4s.... 297 views View 1 Upvoter

What is the next atomic orbital in
the series 1s, 2s, 2p ...

a. 1s 2s 2p 3s 3p c. 1s 2s 2p 3s
3p 4s 4p b. 1s 2s 2p 3s 3p 4s d.

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1s 2s 2p 3s 3p 3d. 1s 2s 2p 3s 3p.

What is the charge on the strontium ion? a. 2- c. 1 b. 1- d. 2.

1-The octet rule states that, in chemical compounds, atoms tend to have _____. a. the electron configuration of a noble gas

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Chemistry test chapter 7 You'll
Remember | Quizlet

An illustration of the shape and relative size of 1s, 2s and 2p orbitals. Click the check boxes to show and hide the atomic orbitals. Explore other atomic

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orbitals. s-orbitals | p-orbitals | d-orbitals. 4.7 (30) How useful was this page? Click on a star to rate it! Submit Rating . Average rating 4.7 / 5.

Atomic Orbitals - shape and

Page 25/41

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relative size of 1s, 2s and 2p ...

Which are impossible? (a) 1s 2 2s
2 3s 2 (b) 1s 2 2p 3 (c) 1s 2 2s 3
2p 5 (d) 1s 2 2s 2 2p 7 (e) 1s 2 2s
2 2p 6 3s 1 (f) 1s 2 2s 2 2p 6 3s 2
3d 1. Buy Find arrow_forward.

Chemistry: Principles and
Reactions. 8th Edition. William L.

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Cengage Learning. ISBN:
9781305079373.

Which of the following electron
configurations (a – f ...

(c) 1s 2s 2p 3s 3p (d) 1s 2s 2p 3s

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3p45 9. (5 points) Please select statements that satisfy the Pauli exclusion principle.

- a) Electron state can hold no more than two electrons.
- b) Electrons with the same state must have opposite spins.
- c) Electron state can hold no more than four electrons.

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Solved: (c) 1s 2s 2p 3s 3p (d) 1s
2s 2p 3s 3p⁴⁵ 9. (5 Poin ...

Mg - $1s^2 2s^2 2p^6 3s^2$ b. F -
 $1s^2 2s^2 2p^5$ c. Si - $1s^2$
 $2s^2 2p^6 3s^2 3p^2$ d. Al-
 $1s^2 2s^2 2p^6 3s^2 3p^1$ e. P

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- $1s^2 2s^2 2p^6 3s^2 3p^3$ f.

Cl - $1s^2 2s^2 2p^6 3s^2 3p^5$

2. Which of the following electron configurations correspond to ground states (lowest energy) and which correspond to excited states? ...

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Chemistry HW6 - sas.upenn.edu
2p 5 b) 1s 2 2s 1 c) 1s 2 2s 2 2p 6
d) 1s 2 s 2 2p 6 3s 2 3p 5 e) 1s 2
2s 2 2p 6 3s 2 3p 6 4s 1 f) 1s 2 2s
2 2p 6 3s 2 3p 6 4s 2 3d 10 4p 6
11. Specify the group of the
periodic table in which each of

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the following elements is found:

- a) [Ne]3s 1 b) [Ne]3s 2 3p 3 c)
[Ne]3s 2 3p 6 d) [Ar]4s 2 3d 8 12.

Arrange the following atoms in
order of ...

2p 5 b 1s 2 2s 1 c 1s 2 2s 2 2p 6

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d 1s 2 s 2 2p 6 3s 2 3p ...

2s shields the atom better than 2p because the s orbitals is much closer and surrounds the nucleus more than the p orbitals, which extend farther out. 3p shields better than 3d, because p orbitals are closer to the nucleus than the

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$3p^2$ \\ \rm (c) Fe &: \rm $1s^2$
 $2s^2 2p^6 3s^2 3p^6 4s^2$
 $3d^6$ \\ \rm (d) Te &: \rm $1s^2$
 $2s^2 2p^6 3s^2 3p \dots$

Using complete subshell notation
($1s^2 2s^2 2p^6$, and so ...

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2 2p 6 3s 1 1s 2 2s 2 2p 6 3s 2 3p
1 1s 2 2s 2 2p 6 3s 2 3p 3 1.

Examine the boarding house diagrams in Model 1. Match each symbol below with its most likely meaning. ____ a. I. Bunk bed for boarders ____ b. II. Manager's code for the number of boarders

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in the house and their room
assignments. ____ c. 1s 2 2s 2 2p
6 3s 1 III. Boarder Sunny Rooms 2
POGIL □ Activities for High School
...

2 2p 6 3s 1 1s 2 2s 2 2p 6 3s 2 3p

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1 1s 2 2s 2 2p 6 3s 2 ...

b. c. d. c. 1s 2s 2p 3s 3p 3d 4s! 1s

2s 2p 3s 3p 4s 3d 1s 2s 2p 3s 3p

3d 1s 2s 2p 3s 3p 3d 1s 2s 2p 3s

3p 45-4d 3. . What is the symbol

of the neutral atom with the

following electron orbital

diagram?

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Solved: 1. Which Of The Following
Electron ... - Chegg.com

There are 3 electrons in 2p
orbitals. Total number of p
electrons in N=3 b) The electronic
configuration of Si : $1s^2 2s^2 2p^2$

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$3s^2 3p^2$ There are two 1s electrons, two 2s electrons, and two 3s electrons. Total number of s electrons in Si = $2 + 2 + 2 = 6$
c) The electronic configuration of S : $1s^2 2s^2 2p^6 3s^2 3p^4$ The 3d orbitals are empty.

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4s C

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