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*#1 (9min) Electron*

*Geometry,*

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smallest  
“uncuttable”  
particle of an  
element is the  
\_\_\_\_\_. 2. Atoms of  
the same element  
have exactly the  
same \_\_\_\_\_. Atoms  
of different  
elements have

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different Table. 3.

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~~3. Atoms & the~~

~~Periodic Table~~

~~Mrs. Sumerix~~

~~Answer~~  
The periodic table helps predict some properties of the elements

compared to each other. Atom size decreases as you move from left to right across the table and increases as you move down

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a column. The energy required to remove an electron from an atom increases as you move from left to right and decreases as you move down a column.

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of Elements 1.

Define Element: An element is a chemical substance that is in the form of an atom

2. Define Atom: An atom is small sub-atomic particles held together by

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fundamental forces

3. Describe changes in the atomic model over time and why those changes were necessitated by experimental evidence. Explain how each discovery lead to the modern ...

~~Atoms and the Per~~

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Name \_\_\_\_\_ The

Scientists A.

Thomson's study of

cathode rays

revealed that

atoms contained

negatively charged

particles called

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Answer  
Rutherford's gold-foil experiment revealed that an atom's positive charge is concentrated in the atom's

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Table Name~~

The position of  
elements on the

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periodic table is directly related to their electron shells. Describe the relationship between group number and number of valence electrons. the group number is how many valence electrons of that element are in each atom.

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Describe the role of  
a coefficient in a  
chemical equation.

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Elements 4) Daily

Warm Ups - Daily

warm ups, "do it

nows" or "bell

activities" serve

the purpose of

engaging students

in the daily

learning, focusing

students for

scientific study and

to practice skills or

learn concepts.

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Mass number. sum  
of the number of  
protons and  
neutrons; number  
of protons +  
number of  
neutrons = mass  
number, number of  
neutrons = mass

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number - number  
of protons (atomic  
number), number  
of protons =  
number of  
electrons IF  
positive and  
negative charges  
cancel, the atom  
Charge = 0 (Stable  
state, non ionic  
state) Isotopes.

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Electron affinity  
increases from left  
to right across the  
periodic table. This  
is caused by the  
decrease in atomic

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radius. As we already explained, moving from left to right across a period, atoms become smaller and smaller as the atomic number increases. This causes the electron to move closer to the nucleus, thus increasing the electron affinity.

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Each row in the  
periodic table is  
called a period and  
the period number  
tells how many  
energy levels  
atoms in that  
period have. As a

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Answer  
result, as you go  
down a group in  
the periodic table  
the sizes...

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size. How does it  
relate to the  
periodic ...~~

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