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The atomic symbol for an ion, it has the mass number, atomic number and the amount of electrons added or subtracted What is the difference between the atomic mass and the atom number The mass number is the protons and neutrons added together and the atomic number is the mass of all isotopes of that element.

Atoms, Isotopes & Ions. Flashcards | Quizlet

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WORKSHEET Atomic # = # of
protons. Mass

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Atomic structure and isotopes answers give the chemical symbol for the following isotopes including the atomic number and mass number. Which of the atoms represented by these symbols are isotopes of each other. Isotopes are atoms of the same element with different mass numbers. The number of protons in an atom is known as the of that atom. Uranium 235 and uranium 238 are considered isotopes of one another. No carbon 12 and carbon 13 are isotopes of carbon but have different mass numbers.

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Isotopes AQA GCSE | Teaching Resources

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mass from isotope data, reading chemical symbols, types of radioactive decay, balancing nuclear equ...

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